

Date Planned : __ / __ / __	Daily Tutorial Sheet - 5	Expected Duration : 90 Min
Actual Date of Attempt : __ / __ / __	Level - 1	Exact Duration : _____

61. The electronic configuration $1s^2 2s^2 2p^6 3s^2 3p^6 3d^{10} 4s^2 4p^6 4d^{10} 5s^2$ is for :
 (A) f-block element (B) d-block element
 (C) p-block element (D) s-block element
62. EN of the element (A) is E_1 and IP is E_2 . Hence according to mulliken scale EA will be: ▶
 (A) $2E_1 - E_2$ (B) $E_1 - E_2$ (C) $E_1 - 2E_2$ (D) $(E_1 + E_2)/2$
63. The element with atomic number 57 belongs to:
 (A) s-block (B) p-block (C) d-block (D) f-block
64. The most metallic of the following elements is:
 (A) Mg (B) Li (C) K (D) Ca
65. Which set expected to show the smallest difference in first ionization energy? ▶
 (A) He, Ne, Ar (B) B, N, O
 (C) Mg, Mg^+ , Mg^{2+} (D) Fe, Co, Ni
66. Successive ionisation energy of an atom is greater than previous one, because: ▶
 (A) $\frac{p}{e}$ ratio increases (B) $\frac{p}{e}$ ratio decreases
 (C) $\frac{p}{e}$ ratio remains constant (D) none of these
67. Correct order of ionisation potential of coinage metals is: ▶
 (A) $Au > Ag > Cu$ (B) $Cu > Ag > Au$
 (C) $Au > Cu > Ag$ (D) $Ag > Cu > Au$
68. On moving from Li to F in the second period, there would be a decrease in: ▶
 (A) non-metallic property (B) atomic radius
 (C) ionization potential (D) electronegativity
69. Which of the following has the maximum value of electronegativity?
 (A) carbon (B) silicon (C) germanium (D) tin
70. Which of the following element has the maximum electron affinity?
 (A) nitrogen (B) oxygen (C) fluorine (D) chlorine
71. A change of Zn to Zn^{2+} is accompanied by a decrease in: ▶
 (A) number of unpaired electrons (B) atomic mass
 (C) atomic number (D) number of shells

Paragraph for Question No. 72 - 74 ▶

Elements with their electronic configurations are given below:

Answer the following questions:

- I. $1s^2 2s^2$ II. $1s^2 2s^2 2p^6$ III. $1s^2 2s^2 2p^6 3s^2$ IV. $1s^2 2s^2 2p^3$
 V. $1s^2 2s^2 2p^5$

72. The element with highest I.E. is:
(A) I (B) II (C) III (D) IV
73. The element with lowest value of electron affinity is:
(A) I (B) II (C) III (D) IV
74. The most ionic compound will be formed between:
(A) I and IV (B) I and V (C) III and IV (D) III and V
75. The number of elements which can be accommodated in the present set-up of the long form of the periodic table is : 
(A) 120 (B) 118 (C) 122 (D) 124